



Western Australian Curriculum

Science

Teacher resources

Years 7–10

Acknowledgement of Country

Kaya. The School Curriculum and Standards Authority (the Authority) acknowledges that our offices are on Whadjuk Noongar boodjar and that we deliver our services on the country of many traditional custodians and language groups throughout Western Australia. The Authority acknowledges the traditional custodians throughout Western Australia and their continuing connection to land, waters and community. We offer our respect to Elders past and present.

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Any resources such as texts, websites and so on that may be referred to in this document are provided as examples of resources that teachers can use to support their learning programs. Their inclusion does not imply that they are mandatory or that they are the only resources relevant to the course.

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Guide to using this resource

This teacher resource is designed to support the implementation of the Western Australian Curriculum: Science in Years 7–10.

Teachers have the flexibility to use this resource according to context and teaching need. They may choose to distribute the whole document to students to use in class and/or for assessments, or to provide extracts relevant to specific year levels and content area/s.

Prefixes of the metric system

| Factor | Prefix | Symbol | Factor | Prefix | Symbol |
|--------|--------|--------|-----------|--------|--------|
| 10^6 | mega | M | 10^{-2} | centi | c |
| 10^3 | kilo | k | 10^{-3} | milli | m |
| | | | 10^{-6} | micro | μ |

Units for base quantities

| Physical quantity | | Unit | |
|-------------------|-----------------|----------------|--------------------|
| Quantity name | Quantity symbol | Unit name | Unit symbol |
| length | l | metre | m |
| mass | m | kilogram | kg |
| time | t | second | s |
| electric current | I | ampere | A |
| temperature | T | degree Celsius | $^{\circ}\text{C}$ |

Units for other common quantities

| Physical quantity | | Unit | |
|-------------------|-----------------|------------------------------|-------------------|
| Quantity name | Quantity symbol | Unit name | Unit symbol |
| volume | V | litre | L |
| force | F | newton | N |
| resistance | R | ohm | Ω |
| voltage | V | volt | V |
| wavelength | λ | metre | m |
| frequency | f | hertz | Hz |
| velocity | v | metres per second | m s^{-1} |
| acceleration | a | metres per second per second | m s^{-2} |

Periodic table of the elements

| | | | | | | | | | | | | | | | | | |
|--------------------------------------|---------------------------------------|--------------------------------------|--|--------------------------------------|---|---------------------------------------|---------------------------------------|---|---|--|--|---------------------------------------|--|--|--|---|---------------------------------------|
| 1 H hydrogen 1.008 | 2 He helium 4.003 | | | | | | | | | | | | | | | | |
| 3 Li lithium 6.94 | 4 Be beryllium 9.012 | | | | | | | | | | | 5 B boron 10.81 | 6 C carbon 12.01 | 7 N nitrogen 14.01 | 8 O oxygen 16.00 | 9 F fluorine 19.00 | 10 Ne neon 20.18 |
| 11 Na sodium 22.99 | 12 Mg magnesium 24.31 | | | | | | | | | | | 13 Al aluminium 26.98 | 14 Si silicon 28.09 | 15 P phosphorus 30.97 | 16 S sulfur 32.06 | 17 Cl chlorine 35.45 | 18 Ar argon 39.95 |
| 19 K potassium 39.10 | 20 Ca calcium 40.08 | 21 Sc scandium 44.96 | 22 Ti titanium 47.87 | 23 V vanadium 50.94 | 24 Cr chromium 52.00 | 25 Mn manganese 54.94 | 26 Fe iron 55.85 | 27 Co cobalt 58.93 | 28 Ni nickel 58.69 | 29 Cu copper 63.55 | 30 Zn zinc 65.38 | 31 Ga gallium 69.72 | 32 Ge germanium 72.63 | 33 As arsenic 74.92 | 34 Se selenium 78.97 | 35 Br bromine 79.90 | 36 Kr krypton 83.80 |
| 37 Rb rubidium 85.47 | 38 Sr strontium 87.62 | 39 Y yttrium 88.91 | 40 Zr zirconium 91.22 | 41 Nb niobium 92.91 | 42 Mo molybdenum 95.95 | 43 Tc technetium | 44 Ru ruthenium 101.1 | 45 Rh rhodium 102.9 | 46 Pd palladium 106.4 | 47 Ag silver 107.9 | 48 Cd cadmium 112.4 | 49 In indium 114.8 | 50 Sn tin 118.7 | 51 Sb antimony 121.8 | 52 Te tellurium 127.6 | 53 I iodine 126.9 | 54 Xe xenon 131.3 |
| 55 Cs caesium 132.9 | 56 Ba barium 137.3 | 57–71 lanthanoids | 72 Hf hafnium 178.5 | 73 Ta tantalum 180.9 | 74 W tungsten 183.8 | 75 Re rhenium 186.2 | 76 Os osmium 190.2 | 77 Ir iridium 192.2 | 78 Pt platinum 195.1 | 79 Au gold 197.0 | 80 Hg mercury 200.6 | 81 Tl thallium 204.4 | 82 Pb lead 207.2 | 83 Bi bismuth 209.0 | 84 Po polonium [209] | 85 At astatine [210] | 86 Rn radon [222] |
| 87 Fr francium [223] | 88 Ra radium [226] | 89–103 actinoids | 104 Rf rutherfordium [267] | 105 Db dubnium [268] | 106 Sg seaborgium [269] | 107 Bh bohrium [270] | 108 Hs hassium [269] | 109 Mt meitnerium [277] | 110 Ds darmstadtium [281] | 111 Rg roentgenium [282] | 112 Cn copernicium [285] | 113 Nh nihonium [286] | 114 Fl flerovium [290] | 115 Mc moscovium [290] | 116 Lv livermorium [293] | 117 Ts tennessine [294] | 118 Og oganeson [290] |

Key:

| |
|------------------------|
| Atomic number |
| Symbol |
| Name |
| Standard atomic weight |

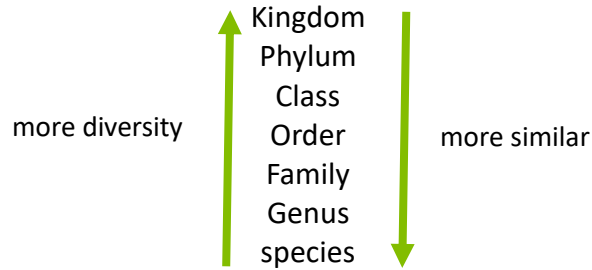
| | | | | | | | | | | | | | | |
|---------------------------------------|-------------------------------------|--|---------------------------------------|--|---------------------------------------|---------------------------------------|--|---------------------------------------|---|---|--------------------------------------|--|---------------------------------------|---|
| 57 La lanthanum 138.9 | 58 Ce cerium 140.1 | 59 Pr praseodymium 140.9 | 60 Nd neodymium 144.2 | 61 Pm promethium [145] | 62 Sm samarium 150.4 | 63 Eu europium 152.0 | 64 Gd gadolinium 157.3 | 65 Tb terbium 158.9 | 66 Dy dysprosium 162.5 | 67 Ho holmium 164.9 | 68 Er erbium 167.3 | 69 Tm thulium 168.9 | 70 Yb ytterbium 173.0 | 71 Lu lutetium 175.0 |
| 89 Ac actinium [227] | 90 Th thorium 232.0 | 91 Pa protactinium 231.0 | 92 U uranium 238.0 | 93 Np neptunium [237] | 94 Pu plutonium [244] | 95 Am americium [243] | 96 Cm curium [247] | 97 Bk berkelium [247] | 98 Cf californium [251] | 99 Es einsteinium [252] | 100 Fm fermium [257] | 101 Md mendelevium [258] | 102 No nobelium [259] | 103 Lr lawrencium [262] |

Data: International Union of Pure and Applied Chemistry, 2022.

Year 7

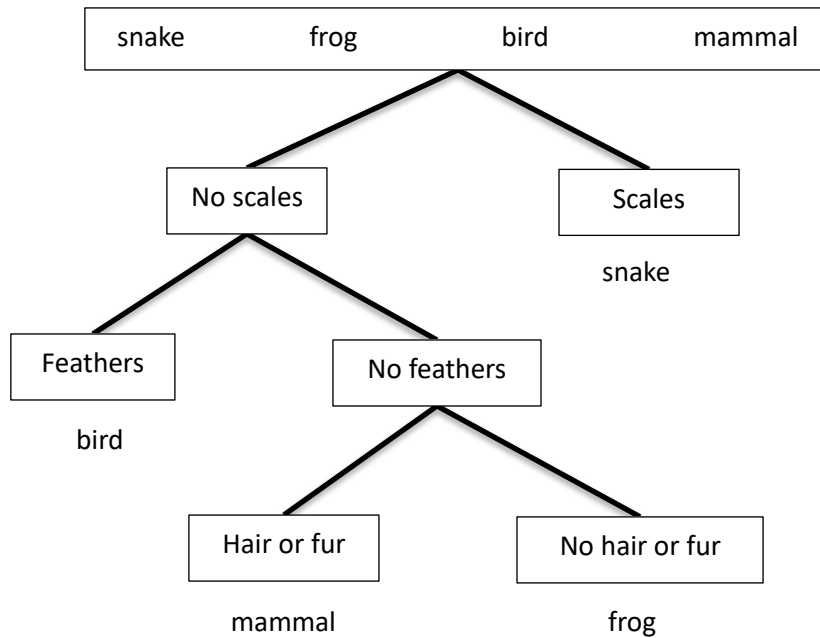
Biological sciences

Hierarchical classification



Dichotomous keys

Flow chart or graphical key



Written or tabular key




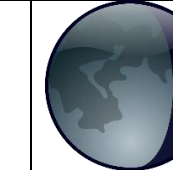



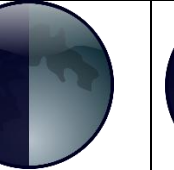
- | | | |
|----|--------------------|---------|
| 1a | Scales ... | snake |
| 1b | No scales ... | go to 2 |
| 2a | Feathers ... | bird |
| 2b | No feathers ... | go to 3 |
| 3a | Hair or fur ... | mammal |
| 3b | No hair or fur ... | frog |

Earth and space sciences




Lunar phases

Lunar phases are the same in the southern and northern hemispheres but the appearance of the moon in each hemisphere is inverted. The phases of the moon in the southern and northern hemispheres are represented below. The white area represents the part of the moon visible from the Earth.

Phases of the moon – southern hemisphere

| New moon | Waxing crescent | First quarter | Waxing gibbous | Full moon | Waning gibbous | Last quarter | Waning crescent |
|---|---|---|---|--|---|---|---|
|  |  |  |  |  |  |  |  |

Phases of the moon – northern hemisphere

| New moon | Waxing crescent | First quarter | Waxing gibbous | Full moon | Waning gibbous | Last quarter | Waning crescent |
|--|--|--|--|---|--|--|--|
|  |  |  |  |  |  |  |  |

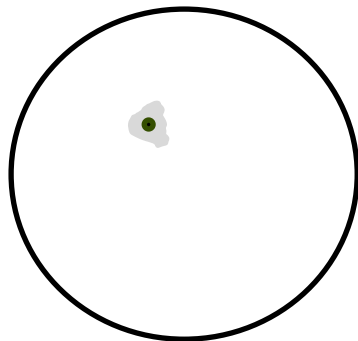
Graphics: De Wulf, n.d.

Year 8

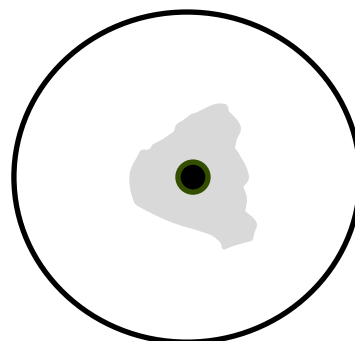
Biological sciences

Calculating magnification

| Ocular lens magnification | Objective lens magnification | Total magnification |
|---------------------------|------------------------------|---------------------|
| ×10 | ×4 | 40 × |
| ×10 | ×10 | 100 × |
| ×10 | ×40 | 400 × |



Low power (40 ×)



High power (400 ×)

Estimating cell size

The following procedure can be used to estimate the size of a cell:

- Use a plastic ruler or mini grid to measure the diameter of the field of view (dFOV) in millimetres (mm).
- Convert the diameter of the field of view to micrometres (μm).
- Estimate the number of cells that will fit across the dFOV.
- Divide the dFOV by the number of cells.

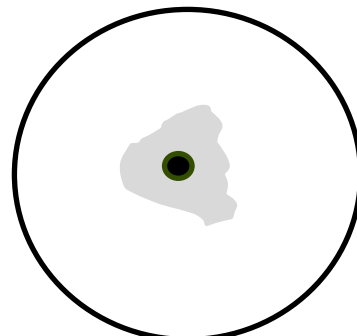
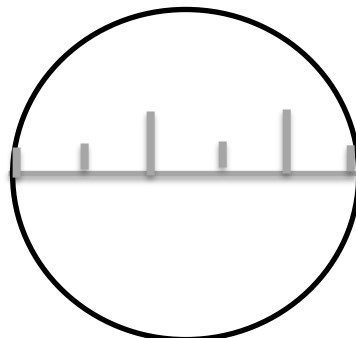
Example:

The magnification is 100 ×

dFOV = 2.5 mm = 2500 μm

Approximately 3 cells fit across the dFOV

thus, size of cell = $\frac{2500}{3} = \sim 833 \mu\text{m}$



Earth and space sciences

Classifying minerals


Lustre

Describes the way minerals reflect light from their surface. Lustre can be classified as:

- Glassy – look like glass
- Pearly – like the surface of a pearl
- Metallic – shiny like the surface of a metal
- Dull – no reflection
- Earth – has a surface texture like soil.

Mohs hardness scale

The hardness of a mineral is its resistance to being scratched. Mohs hardness scale can be used to classify hardness.

| Rank position on Mohs scale softest | Standard reference mineral | Approximate equivalents |
|--|-------------------------------|-------------------------|
|  | 1 talc | |
| | 2 gypsum | fingernail |
| | 3 calcite | copper coin |
| | 4 fluorite | |
| | 5 apatite | window glass |
| | 6 feldspar | steel knife blade |
| | 7 quartz | |
| | 8 topaz | masonry drill bit |
| | 9 corundum | |
| | 10 diamond | |

Density

$$\text{density} = \frac{\text{mass}}{\text{volume}}$$

Cleavage

Some minerals cleave or split along planes due to their structure. Broken surfaces of mineral specimens can be classified as:

- Perfect – always flat, mirror like
- Good – sometimes flat, mirror-like surfaces
- Poor – occasionally flat, mirror-like surfaces
- None – no cleavage.

Chemical sciences

Symbols for common elements

| Element name | Element symbol | Element name | Element symbol |
|--------------|----------------|--------------|----------------|
| Hydrogen | H | Chromium | Cr |
| Helium | He | Iron | Fe |
| Lithium | Li | Cobalt | Co |
| Beryllium | Be | Nickel | Ni |
| Boron | B | Copper | Cu |
| Carbon | C | Zinc | Zn |
| Nitrogen | N | Bromine | Br |
| Oxygen | O | Silver | Ag |
| Fluorine | F | Iodine | I |
| Neon | Ne | Tin | Sn |
| Sodium | Na | Barium | Ba |
| Magnesium | Mg | Platinum | Pt |
| Aluminium | Al | Gold | Au |
| Silicon | Si | Mercury | Hg |
| Phosphorus | P | Lead | Pb |
| Sulfur | S | Uranium | U |
| Chlorine | Cl | | |
| Argon | Ar | | |
| Potassium | K | | |
| Calcium | Ca | | |

Names and formulae for the molecules of common elements



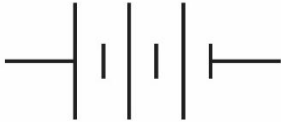




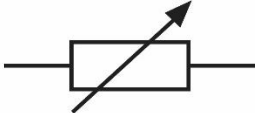


| Name | Formula |
|----------|-----------------|
| Hydrogen | H ₂ |
| Nitrogen | N ₂ |
| Oxygen | O ₂ |
| Fluorine | F ₂ |
| Chlorine | Cl ₂ |
| Iodine | I ₂ |

Names and formulae of common compounds

| Name | Formula |
|--------------------|---|
| Water | H ₂ O |
| Carbon dioxide | CO ₂ |
| Sulfur dioxide | SO ₂ |
| Ammonia | NH ₃ |
| Methane | CH ₄ |
| Glucose | C ₆ H ₁₂ O ₆ |
| Acetic acid | CH ₃ COOH |
| Hydrochloric acid | HCl |
| Sulfuric acid | H ₂ SO ₄ |
| Sodium chloride | NaCl |
| Copper(II) sulfate | CuSO ₄ |

Physical sciences





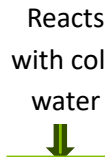
Electrical circuit diagram symbols

| Name | Symbol | Name | Symbol |
|----------------------|---|-------------------|---|
| electrochemical cell |  | conductor wire |  |
| battery |  | light globe |  |
| open switch |  | resistor |  |
| closed switch |  | variable resistor |  |
| ammeter |  | voltmeter |  |

Year 9

Chemical sciences

Activity series of selected metals

| | Metal | Symbol | | | | |
|---|-----------|--------|---|--|---|--|
| Most reactive  Least reactive | Lithium | Li |  Reacts with oxygen |  Reacts with acids |  Reacts with boiling water |  Reacts with cold water |
| | Caesium | Cs | | | | |
| | Rubidium | Rb | | | | |
| | Potassium | K | | | | |
| | Sodium | Na | | | | |
| | Barium | Ba | | | | |
| | Strontium | Sr | | | | |
| | Calcium | Ca | | | | |
| | Magnesium | Mg | | | | |
| | Aluminium | Al | | | | |
| | Titanium | Ti | | | | |
| | Manganese | Mn | | | | |
| | Zinc | Zn | | | | |
| | Chromium | Cr | | | | |
| | Iron | Fe | | | | |
| | Cobalt | Co | | | | |
| | Nickel | Ni | | | | |
| | Tin | Sn | | | | |
| | Lead | Pb | | | | |
| | Copper | Cu | | | | |
| Tungsten | W | | | | | |
| Mercury | Hg | | | | | |
| Silver | Ag | | | | | |
| Gold | Au | | | | | |
| Platinum | Pt | | | | | |

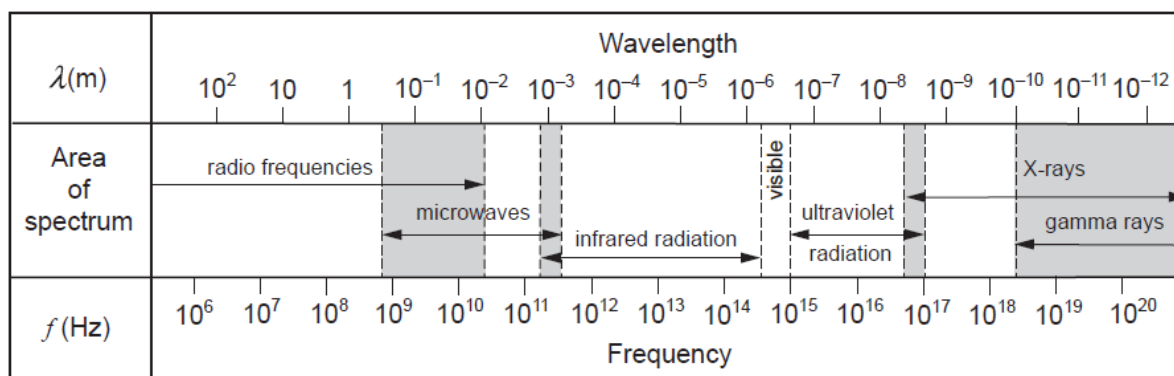
Common monatomic and polyatomic ions

| Ion name | Formula |
|------------|------------------|
| ammonium | NH_4^+ |
| copper(I) | Cu^+ |
| hydrogen | H^+ |
| lithium | Li^+ |
| potassium | K^+ |
| silver | Ag^+ |
| sodium | Na^+ |
| barium | Ba^{2+} |
| calcium | Ca^{2+} |
| copper(II) | Cu^{2+} |
| iron(II) | Fe^{2+} |
| lead(II) | Pb^{2+} |
| magnesium | Mg^{2+} |
| zinc | Zn^{2+} |
| aluminium | Al^{3+} |
| iron(III) | Fe^{3+} |

| Ion name | Formula |
|-------------------|--------------------|
| bromide | Br^- |
| chloride | Cl^- |
| fluoride | F^- |
| hydrogencarbonate | HCO_3^- |
| hydroxide | OH^- |
| iodide | I^- |
| nitrate | NO_3^- |
| carbonate | CO_3^{2-} |
| oxide | O^{2-} |
| sulfate | SO_4^{2-} |
| sulfide | S^{2-} |
| nitride | N^{3-} |
| phosphate | PO_4^{3-} |

Physical sciences

Electromagnetic spectrum



Note: shaded areas represent regions of overlap.

Year 10

Biological sciences

DNA nitrogenous bases

| Nitrogenous base | Abbreviation |
|------------------|--------------|
| adenine | A |
| cytosine | C |
| guanine | G |
| thymine | T |

Nitrogenous base pairs:

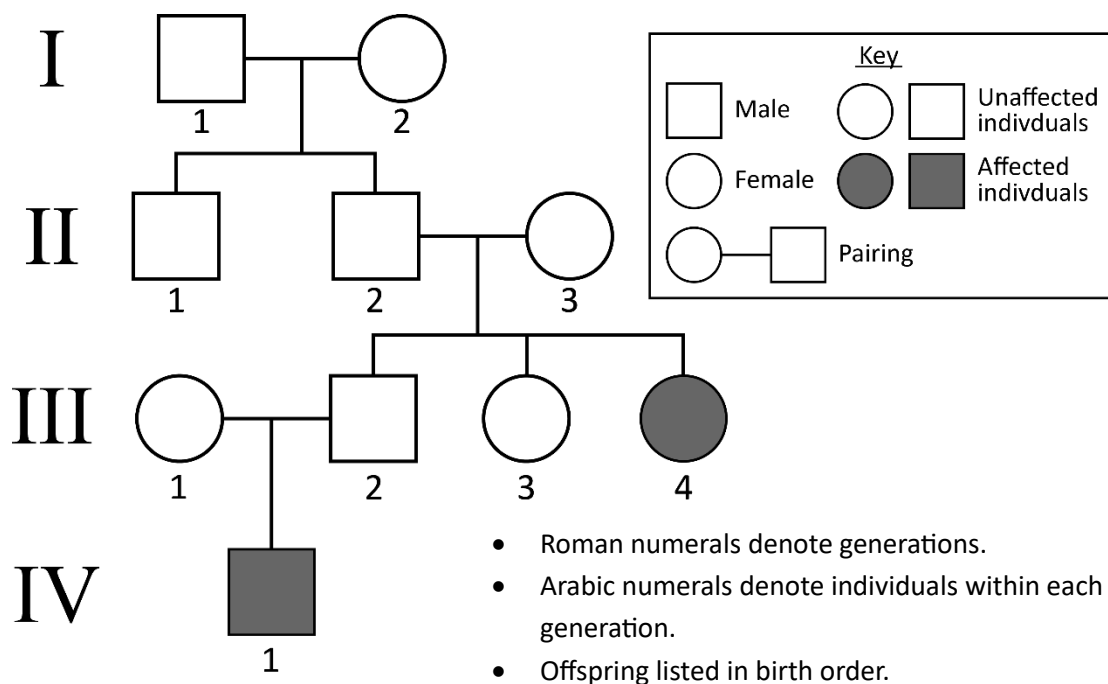
- Adenine and thymine
- Guanine and cytosine.

Representing alleles

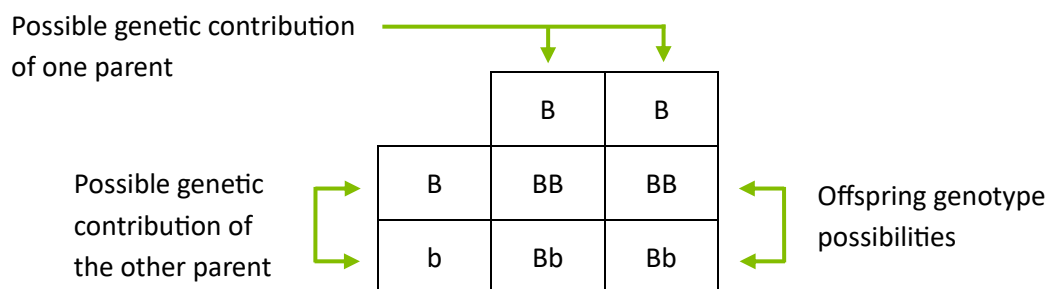
Upper case – dominant (B)
Lower case – recessive (b)

- Representation of autosomal alleles as genotype – BB, Bb and bb
- Representation of sex-linked alleles as genotype – X^BX^b and X^bY , where X represents the female chromosome and Y the male chromosome.

Pedigree chart



Punnett square



This Punnett square shows the possible genotype combinations for a dominant/recessive pattern of inheritance; this would mean all offspring would display the dominant phenotype.

Chemical sciences

Common monatomic and polyatomic ions

| Ion name | Formula |
|---------------|------------------|
| ammonium | NH_4^+ |
| copper(I) | Cu^+ |
| hydrogen | H^+ |
| lithium | Li^+ |
| potassium | K^+ |
| silver | Ag^+ |
| sodium | Na^+ |
| barium | Ba^{2+} |
| calcium | Ca^{2+} |
| cobalt(II) | Co^{2+} |
| copper(II) | Cu^{2+} |
| iron(II) | Fe^{2+} |
| lead(II) | Pb^{2+} |
| magnesium | Mg^{2+} |
| manganese(II) | Mn^{2+} |
| nickel(II) | Ni^{2+} |
| strontium | Sr^{2+} |
| zinc | Zn^{2+} |
| aluminium | Al^{3+} |
| chromium(III) | Cr^{3+} |
| iron(III) | Fe^{3+} |

| Ion name | Formula |
|---------------------|---------------------------|
| bromide | Br^- |
| chloride | Cl^- |
| acetate (ethanoate) | CH_3COO^- |
| fluoride | F^- |
| hydrogencarbonate | HCO_3^- |
| hydroxide | OH^- |
| iodide | I^- |
| nitrate | NO_3^- |
| carbonate | CO_3^{2-} |
| oxide | O^{2-} |
| sulfate | SO_4^{2-} |
| sulfide | S^{2-} |
| nitride | N^{3-} |
| phosphate | PO_4^{3-} |

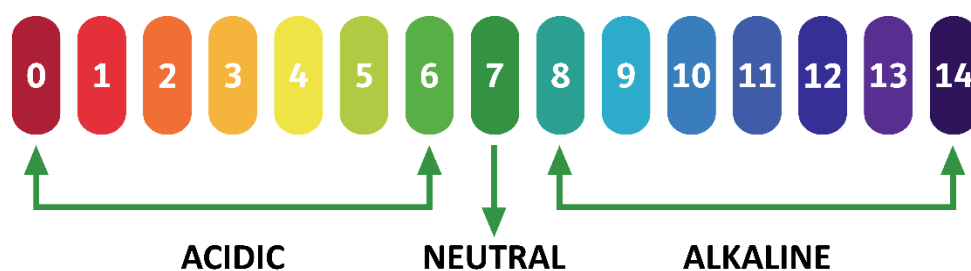
Common acids and bases

| Acid name | Formula |
|-------------------|--------------------------|
| acetic acid | CH_3COOH |
| hydrochloric acid | HCl |
| nitric acid | HNO_3 |
| carbonic acid | H_2CO_3 |
| sulfuric acid | H_2SO_4 |
| phosphoric acid | H_3PO_4 |

| Base name | Formula |
|---------------------|-------------------|
| ammonia* | NH_3 |
| sodium hydroxide | NaOH |
| potassium hydroxide | KOH |
| calcium hydroxide | Ca(OH)_2 |
| magnesium hydroxide | Mg(OH)_2 |

*In aqueous solution ammonia (NH_3) produces NH_4^+ and OH^- ions

pH scale



Solubility rules for ionic solids in water

Soluble in water

| Soluble | Insoluble |
|----------------|---|
| Most chlorides | AgCl |
| Most bromides | AgBr |
| Most iodides | AgI , CuI , PbI_2 |
| All nitrates | No exceptions |
| All ethanoates | |
| Most sulfates | SrSO_4 , BaSO_4 , PbSO_4 |

Insoluble in water

| Insoluble | Soluble |
|-----------------|--|
| Most hydroxides | NaOH , KOH , Ba(OH)_2 , NH_4OH^* , AgOH^{**} |
| Most carbonates | Na_2CO_3 , K_2CO_3 , $(\text{NH}_4)_2\text{CO}_3$ |
| Most phosphates | Na_2PO_4 , K_2PO_4 , $(\text{NH}_4)_3\text{PO}_4$ |
| Most sulfides | Na_2S , K_2S , $(\text{NH}_4)_2\text{S}$ |

* NH_3 dissolves in water to form both $\text{NH}_3(\text{aq})$ and the ions $\text{NH}_4^+(\text{aq})$ and $\text{OH}^-(\text{aq})$

** $\text{Ag}^+(\text{aq})$ reacts with $\text{OH}^-(\text{aq})$ to form insoluble Ag_2O

Soluble = more than 0.1 mole dissolves per litre
 Insoluble = less than 0.01 mole dissolves per litre

General chemical reactions for predicting products

Acid + metal hydroxide (base) → salt + water

Acid + metal oxide (base) → salt + water

Acid + metal → salt + hydrogen gas

Acid + carbonate → salt + water + carbon dioxide gas

Acid + hydrogencarbonate → salt + water + carbon dioxide gas

Physical sciences

Motion, forces and energy

Time difference

$$\Delta t = t_f - t_i$$

Displacement

$$s = \Delta x = x_f - x_i$$

Average velocity

$$v = \frac{\Delta x}{\Delta t} = \frac{x_f - x_i}{t_f - t_i}$$

Equations of motion

$$a = \frac{\Delta v}{\Delta t} = \frac{v_f - v_i}{t_f - t_i} \quad v_f = v_i + a\Delta t$$

$$s = v_i\Delta t + \frac{1}{2}a\Delta t^2$$

Force

$$\Sigma F = ma$$

Weight force

$$F_{\text{weight}} = mg$$

Mechanical energy

Kinetic energy

$$E_k = \frac{1}{2}mv^2$$

Gravitational potential energy

$$E_p = mg\Delta h$$

Conservation of energy

$$\Sigma E_i = \Sigma E_f$$

For an object moving up or down in a gravitational field

$$\Delta E_k = \Delta E_p$$

Work done

$$W = \Delta E$$

Key to symbols

t = time in seconds (s)

s = displacement in metres (m)

x = distance in metres (m)

v = velocity in metres per second (m s^{-1})

a = acceleration in metres per second per second (m s^{-2})

F = force in newtons (N)

m = mass in kilograms (kg)

g = acceleration due to gravity (Mean acceleration due to gravity on Earth = 9.80 m s^{-2})

h = height of an object above a reference point in a gravitational field

The subscripts i and f indicate the final and initial states, respectively. For example, v_i represents the initial velocity of an object and v_f represents the object's final velocity.

Acknowledgements

| | |
|----------------|---|
| Periodic table | Data from: International Union of Pure and Applied Chemistry. (2022). <i>IUPAC Periodic Table of the Elements</i> . Retrieved June, 2025, from https://iupac.org/what-we-do/periodic-table-of-elements/ |
| Lunar phases | <p>Southern hemisphere graphics from, and northern hemisphere graphics adapted from:</p> <p>De Wulf, C. (n.d.). <i>01 Nomoon Openclipart c. 01</i> [Graphic]. Retrieved December, 2025, from https://openclipart.org/detail/119851/01-nomoon-openclipart-c-01</p> <p>De Wulf, C. (n.d.). <i>08 Waningcrescent Openc 01</i> [Graphic]. Retrieved December, 2025, from https://openclipart.org/detail/119809/08-waningcrescent-openc-01</p> <p>De Wulf, C. (n.d.). <i>07 Lastquarter Openclip 01</i> [Graphic]. Retrieved December, 2025, from https://openclipart.org/detail/119815/07-lastquarter-openclip-01</p> <p>De Wulf, C. (n.d.). <i>06 Waninggibbous Opencil 01</i> [Graphic]. Retrieved December, 2025, from https://openclipart.org/detail/119821/06-waninggibbous-opencil-01</p> <p>De Wulf, C. (n.d.). <i>05 Fullmoon Openclipart 01</i> [Graphic]. Retrieved December, 2025, from https://openclipart.org/detail/119845/05-fullmoon-openclipart-01</p> <p>De Wulf, C. (n.d.). <i>04 Waxinggibbous Opencil 01</i> [Graphic]. Retrieved December, 2025, from https://openclipart.org/detail/119827/04-waxinggibbous-opencil-01</p> <p>De Wulf, C. (n.d.). <i>03 Firstquarter Opencil 01</i> [Graphic]. Retrieved December, 2025, from https://openclipart.org/detail/119833/03-firstquarter-opencil-01</p> <p>De Wulf, C. (n.d.). <i>02 Waxingcrescent Openc 01</i> [Graphic]. Retrieved December, 2025, from https://openclipart.org/detail/119839/02-waxingcrescent-openc-01</p> |